

Ag(nic)₂ (nic = Nicotinate): A Spin-Canted Quasi-2D Antiferromagnet Composed of Square-Planar $S = 1/2$ Ag^{II} Ions

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Supporting Information

ABSTRACT: Square-planar $S = 1/2$ Ag^{II} ions in polymeric Ag(nic)₂ are linked by bridging nic monoanions to yield 2D corrugated sheets. Long-range magnetic order occurs below $T_N = 11.8(2)$ K due to interlayer couplings that are estimated to be about 30 times weaker than the intralayer exchange interaction.

Pyridinecarboxylate ligands, and especially nicotinate (nic), have been used extensively in the self-assembly of 1-, 2-, and 3D polymeric architectures. This is because nic can serve as an effective bridging function (through pyridine N and carboxylate O atoms) between transition-metal cations.^{1–3} Structurally characterized examples include, but are not limited to, $M(\text{nic})_2$ ($M = \text{Mn},^{1a} \text{Fe},^{1b} \text{Co},^{1c}$ and Cu^{1d}), $\text{Ni}_2(\text{nic})_4(\text{H}_2\text{O})_2^{2a}$, $\text{Co}(\text{nic})_2(\text{H}_2\text{O})_2^{2b}$, $\text{Mn}(\text{nic})_2(\text{H}_2\text{O})_2^{2c}$, $\text{Ni}(\text{nic})_2(\text{H}_2\text{O})_4^{2d}$ and $\text{Ag}(\text{nic})(\text{nicH})^3$ among others. For $M(\text{nic})_2$, $M = \text{Mn}, \text{Fe}$, and Co are characterized by six-coordinate MO_4N_2 centers connected to other sites via either μ - or μ_3 -coordination or a combination thereof. $\text{Cu}(\text{nic})_2$ has distorted five-coordinate CuO_3N_2 units, with each nic anion being ligated through the N atom and both O atoms to three different Cu ions.^{1d} The few reports on their magnetic properties indicated only weak exchange interactions and the absence of long-range magnetic ordering (LRO) above 2 K.

Although the $S = 1/2$ Ag^{II} ion is generally oxygen-sensitive, Banerjee and Ray reported the synthesis of $\text{Ag}(\text{nic})_2$ in 1956 and noted that it could remain stable for several months so long as the powder was kept dry.⁴ Early powder X-ray diffraction studies on $\text{Ag}(\text{nic})_2$ suggested a tetragonal symmetry, although a detailed structural analysis was not carried out.⁵ Adopting a slightly modified preparation method (see SI), samples were prepared at low temperature and stored in a refrigerator; unlike $\text{Ag}(\text{pyz})_2(\text{S}_2\text{O}_8)^6$, we ultimately found $\text{Ag}(\text{nic})_2$ to be air-stable at room temperature.

Herein, we report on the crystal structure and magnetism of $\text{Ag}(\text{nic})_2$. Using X-rays at the National Synchrotron Light Source (Brookhaven National Laboratory), the crystal structure of $\text{Ag}(\text{nic})_2$ was solved from powder X-ray diffraction data^{7,8} (see Figure S1). The Ag^{II} ion has square-planar D_{2h} symmetry and occupies an inversion center. The coordination sphere

consists of two O atoms [$\text{Ag1-O1} = 2.18(1)$ Å] and two N atoms [$\text{Ag1-N1} = 2.195(5)$ Å] that belong to four different nic anions. A second carboxylate O atom lies 2.95 Å away and is only weakly interacting [$\sum \text{vdw} = 1.72 (\text{Ag}) + 1.52 (\text{O}) = 3.24$ Å]. The AgO_2N_2 plane exhibits a nearly ideal O1–Ag1–N1 bond angle of $89.7(2)^\circ$, and the Ag1–O1–C6 bond angle formed upon coordination of the carboxylate O1 atom is $108.9(6)^\circ$. The carboxylate moiety makes a torsion angle of 12.4° relative to the pyridine ring.

$\text{Ag}(\text{nic})_2$ consists of extended 2D polymeric networks (Figure 1). Every AgO_2N_2 unit is connected to four others

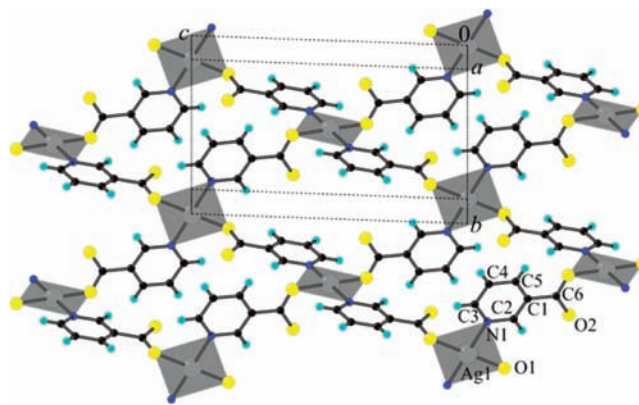


Figure 1. Portion of a 2D polymeric sheet as found in $\text{Ag}(\text{nic})_2$.

via bridging nic anions to form a rhombically distorted square array of equidistant Ag^{II} ions ($\text{Ag}\cdots\text{Ag} = 8.205$ Å). The proximity of the N atom and carboxylate moiety on the nic anion leads to corrugated sheets such that every other AgO_2N_2 unit adopts the same orientation. The sheets in $\text{Cu}(\text{nic})_2$ are also pleated but not in the same way as $\text{Ag}(\text{nic})_2$ because of the differences in nic and Cu^{II} ion coordination.

Adjacent 2D sheets pack *in*-registry along the $[\bar{1}01]$ direction (Figure S2) whereby AgO_2N_2 planes stack directly above and below other AgO_2N_2 units of the same spatial orientation. The

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closest interlayer Ag...Ag spacing is 5.016 Å. The structure of Ag(nic)₂ represents a new topology among M(nic)₂ coordination polymers.

The magnetic susceptibility, $\chi(T)$, for Ag(nic)₂ was measured in a 0.1 T direct-current (dc) magnetic field (Figure 2a). χ

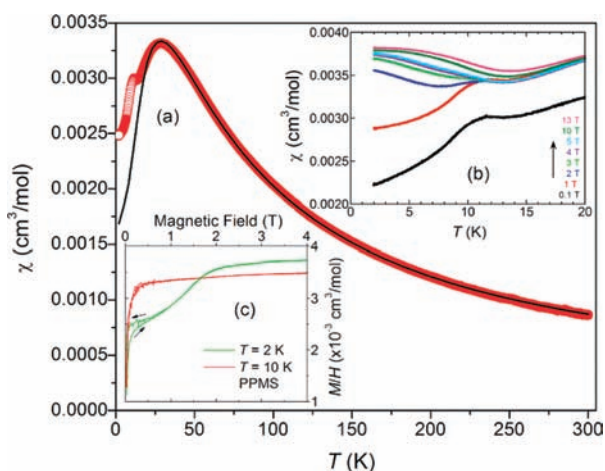


Figure 2. (a) Magnetic susceptibility data for Ag(nic)₂ measured at $H_{dc} = 0.1$ T. The solid line is a theoretical fit to the data as described in the text. (b) $M(T)$ between 2 and 20 K for several H . (c) $M(H)$ obtained at 1.5 and 4.1 K for pulsed fields up to 60 T.

gradually increases as the temperature is lowered, passing through a broad maximum at $T_{max} = 28.7$ K. For low-dimensional magnetic systems, the presence of a broad maximum typically signifies short-range spin correlations due to the largest exchange interaction, in this case, between $S = 1/2$ Ag^{II} ions within layers. Below T_{max} , χ continues to decrease until a second weak peak occurs at 11.5 K, which we ascribe to 3D long-range magnetic order (T_N), with the energy scale set by a weaker interlayer interaction. It is interesting to note that such a feature is completely absent in Ag(py_z)₂(S₂O₈)⁶ and may originate from a larger interlayer coupling and/or canting of the magnetic moments in Ag(nic)₂. The notion of spin canting in Ag(nic)₂ is supported by the field dependence of $M(T)$ in the vicinity of the 11.5 K feature (Figure 2b), whereas the 28 K feature remains broad and does not shift with the field.

An estimate of the relative strength of magnetic interactions in Ag(nic)₂ can be gleaned from a fit ($100 \leq T \leq 300$ K) of the reciprocal magnetic susceptibility, $1/\chi(T)$ (not shown), to a Curie–Weiss law, where $1/\chi = 8(T - \theta)/N\mu_B^2 g^2 S(S + 1)$. The fitted parameters were $g = 2.12(1)$ and $\theta = -46.0(1)$ K, where $\theta < 0$ is consistent with antiferromagnetic (AFM) coupling between $S = 1/2$ Ag^{II} sites. $\chi(T)T$ (not shown) strongly decreases over the whole T range, which is consistent with strong AFM interactions.

The intrinsic two-dimensionality of the crystal structure, as well as the plausible Heisenberg behavior of the Ag^{II} ion and the negative Weiss constant, calls for an AFM Heisenberg model based on a square lattice. Accordingly, the $\chi(T)$ data presented in Figure 2a were fitted⁹ over the 16–300 K temperature range (i.e., above T_N) to yield satisfactory agreement for $g = 2.10(1)$ and $J/k_B = 30.3(1)$ K based on the spin Hamiltonian $\hat{H} = J \sum_i S_i \cdot S_j$, where $J > 0$ indicates AFM coupling. The resulting J value is in line with the temperature of the peak susceptibility, $T_{max} = 0.936J = 28.4$ K, obtained from high-temperature series expansions.¹⁰ Ag(nic)₂ displays much

stronger AFM coupling than Cu(nic)₂ ($J = 12.8$ K), despite the presence of shorter Cu–O–C–O–Cu exchange paths in the latter.^{1d} Together, the nonlinear Ag–nic–Ag pathways and longer Ag...Ag separations in Ag(nic)₂ afford smaller AFM coupling relative to Ag(py_z)₂(S₂O₈) where $J = 53$ K.⁶

The bulk magnetic behavior of Ag(nic)₂ can be rationalized by considering the likely exchange pathways. In this system, the magnetic $d_{x^2-y^2}$ orbital of Ag^{II} overlaps with the lone pairs of electrons belonging to N1 and O1 of the bridging nic anion. The exchange interaction is mediated between adjacent Ag^{II} ions via the σ -bonding network provided by the bridging ligands. The d_{z^2} orbital is spin-paired, lies orthogonal to the AgO₂N₂ plane, and cannot provide a suitable exchange pathway. Appreciable tilting of the AgO₂N₂ planes with respect to nearest-neighbors likely induces a staggered g -tensor, which affords a spin-canted ground state, as suggested by $M(H, T)$ data.

A strong magnetic coupling between Ag^{II} ions is evidenced by the pulsed-field magnetization data (Figure S3), whereby (i) a small moment was found and (ii) a very slightly concave $M(H)$ curve, typical of a quasi-2D $S = 1/2$ Heisenberg antiferromagnet, was observed.¹¹ The latter characteristic persists to the highest available field, suggesting a saturation field (B_c) well above 60 T. From previous work, we found that the simple expression $gB_c/J \approx 6.03$ T/K¹¹ may be used to estimate B_c of a quasi-2D $S = 1/2$ antiferromagnet based on the extracted g and J values obtained from the fit of $\chi(T)$. The extracted J value, known to describe Cu^{II} salts rather well, predicts $B_c \approx 87$ T for Ag(nic)₂. The $M(H)$ data suggest a higher critical field, perhaps due to an additional energy scale, or a significantly enhanced concavity of $M(H)$ above 60 T, which implies a lower magnetic dimensionality. Figure 2c clearly shows that by ~ 1 and 2 T, respectively, the hysteresis between up/downfield sweeps and a spontaneous magnetic moment becomes quenched, whereas M/H grows nearly linear for larger H as a result of dominant in-plane AFM correlations.

Zero-field (ZF) muon-spin-rotation experiments were performed to confirm the suspected long-range magnetic order in Ag(nic)₂. Example spectra are shown in the inset of Figure 3. For $T > 12$ K, we observe monotonic relaxation,

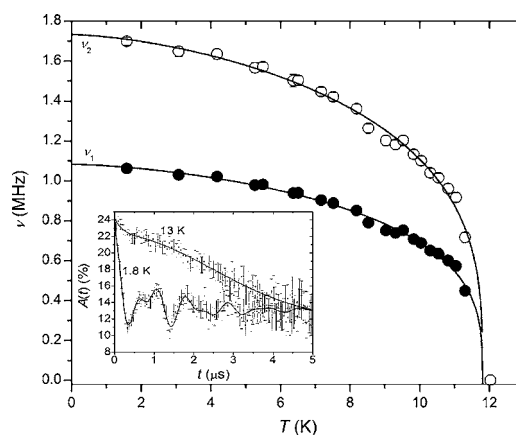


Figure 3. (Inset) ZF μ^+ SR data measured at 1.8 and 13 K for Ag(nic)₂. Oscillations were observed for $T < 12$ K. (Main) T evolution of the muon precession frequencies in Ag(nic)₂. For both plots, the solid lines denote theoretical fits, as described in the text and SI.

typical of $T > T_N$ behavior in materials of this type.^{6,12} Below the AFM transition at T_N , we see oscillations in the time

dependence of the muon polarization [the “asymmetry” $A(t)^{13}$], which are characteristic of a quasi-static local magnetic field at the muon stopping site. This provides very strong evidence for the existence of long-range magnetic order.¹⁴

The muon precession frequencies (ν), of which two exist in $\text{Ag}(\text{nic})_2$, are proportional to the magnetic order parameter. Two muon precession frequencies are typically observed in molecular materials of this type. However, without a magnetic structure, it is difficult to reliably determine the muon stopping sites. In a recent paper,¹⁴ we demonstrated that there are several plausible muon sites in one class of coordination polymers, some of which do not lead to measurable oscillations. Given that the role of the μSR measurements in this study was to determine the presence of LRO rather than an exhaustive analysis of the muon response, we will not speculate on their positions in this structure.

The T dependence of ν is shown in the main plot of Figure 3. The behavior may be parametrized by a fit to the phenomenological expression $\nu_i(T) = \nu_i(0) [1 - (T/T_N)^\alpha]^\beta$, from which we extract $\nu_1(0) = 1.08(1)$ MHz, $\nu_2(0) = 1.73(2)$ MHz, $T_N = 11.8(2)$ K, $\alpha = 1.5(3)$, and $\beta = 0.30(4)$. An increase in the nonoscillatory component near T_N makes it difficult to come to firm conclusions about the critical behavior of $\text{Ag}(\text{nic})_2$.

ZF $\mu^+\text{SR}$ measurements were made up to 50 K, and there is no evidence for changes in the magnetic behavior around 28 K, the temperature at which $\chi(T)$ exhibits a maximum. In fact, we would not expect any; above T_N in these materials, the electronic moments typically fluctuate at a very rapid rate compared to the muon response time (determined by the muon gyromagnetic ratio) and are therefore motionally narrowed from the spectra. The muon relaxation above T_N is then caused by disordered nuclear moments. We therefore do not generally observe any feature in the muon spectra around T_{max} of $\chi(T)$, which is temperature-independent above T_N .

By combining the known T_N and J values, we estimated the interlayer magnetic interaction (J_\perp) using the 2D model obtained by Yasuda et al.¹⁵ From the equation $\ln(J_\perp/J) = b - 4\pi\rho_s/T_N$, where $b = 2.43$ and $\rho_s = 0.183J$ for $S = 1/2$, we calculate $|J_\perp/J| = 0.033$ or $|J_\perp| = 0.99$ K. The small J_\perp is sufficient to facilitate the observed LRO that occurs.

In conclusion, $\text{Ag}(\text{nic})_2$ has a unique 2D corrugated structure composed of alternately tilted square-planar AgO_2N_2 units that consist of Ag^{II} ($S = 1/2$) centers. Magnetic data suggest that it is a quasi-2D Heisenberg antiferromagnet likely to exhibit a spin-canted ground state below $T_N = 11.8(2)$ K. A large AFM interaction occurs within the 2D layers, although the coupling is much weaker between them. To the best of our knowledge, $\text{Ag}(\text{nic})_2$ is the first reported molecule-based Ag^{II} -containing spin-canted AFM.

■ ASSOCIATED CONTENT

■ Supporting Information

Synthesis, characterization details, layer packing diagram, high-field $M(H)$ plot, and CIF for $\text{Ag}(\text{nic})_2$. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Notes

The authors declare no competing financial interest.

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■ REFERENCES

- (1) (a) Lin, W.; Chapman, M. E.; Wang, Z.; Lee, G. T. *Inorg. Chem.* **2000**, *39*, 4169. (b) Ng, S. W. *Acta Crystallogr.* **2008**, *E64*, m728. (c) Feng, W.-J.; Zhou, G.-P.; Zheng, X.-F.; Liu, Y.-G.; Xu, Y. *Acta Crystallogr.* **2006**, *E62*, m2033. (d) Chapman, M. E.; Ayyappan, P.; Foxman, B. M.; Lee, G. T.; Lin, W. *Cryst. Growth Des.* **2001**, *1*, 159.
- (2) (a) Wu, C.-D.; Lu, C.-Z.; Zhuang, H.-H.; Huang, J.-S. *Z. Anorg. Allg. Chem.* **2003**, *629*, 693. (b) Yeh, C.-W.; Suen, M.-C.; Hu, H.-L.; Chen, J.-D.; Wang, J.-C. *Polyhedron* **2004**, *23*, 1947. (c) Hao, X.; Wei, Y.-G.; Liu, Q.; Zhang, S.-W. *Acta Crystallogr.* **2000**, *C56*, 296. (d) Batten, S. R.; Harris, A. R. *Acta Crystallogr.* **2001**, *E57*, m9.
- (3) Kall, P. O.; Grins, J.; Fahlman, M.; Soderlind, F. *Polyhedron* **2001**, *20*, 2747.
- (4) Banerjee, B.; Ray, P. J. *Ind. Chem. Soc.* **1956**, *33*, 503.
- (5) Chackraburty, D. M. *Acta Crystallogr.* **1957**, *10*, 128.
- (6) Manson, J. L.; Stone, K. H.; Southerland, H. I.; Lancaster, T.; Steele, A. J.; Blundell, S. J.; Pratt, F. L.; Baker, P. J.; McDonald, R. D.; Sengupta, P.; Singleton, J.; Goddard, P. A.; Lee, C.; Whangbo, M.-H.; Warter, M. L.; Mielke, C. H.; Stephens, P. W. *J. Am. Chem. Soc.* **2009**, *131*, 4590.
- (7) (a) Bruker AXS. *TOPAS V4: General profile and structure analysis software for powder diffraction data User's Manual*; Bruker AXS: Karlsruhe, Germany, 2005. (b) Coelho, A. A. *J. Appl. Crystallogr.* **2003**, *36*, 86. (c) *TOPAS-Academic* is available at www.topas-academic.net.
- (8) Structural data for $\text{Ag}(\text{nic})_2$: empirical formula $\text{C}_{12}\text{H}_8\text{AgN}_2\text{O}_4$, $M_r = 352.07$, space group $P2_1/n$, $T = 300$ K, $a = 5.0157(1)$ Å, $b = 8.0371(2)$ Å, $c = 14.1498(4)$ Å, $\beta = 98.373(3)^\circ$, $V = 564.33(2)$ Å³, $Z = 2$, $\rho = 2.072$ g cm⁻³; $\mu = 1.697$ mm⁻¹; $R_{\text{wp}} = 0.05568$, $R_{\text{exp}} = 0.04615$. Further details of the structure solution and refinement are provided in the Supporting Information and CIF.
- (9) Woodward, F. M.; Albrecht, A. S.; Wynn, C. M.; Landee, C. P.; Turnbull, M. M. *Phys. Rev. B* **2002**, *65*, 144412.
- (10) Lines, M. E. *J. Phys. Chem. Solids* **1970**, *31*, 101.
- (11) Goddard, P. A.; Singleton, J.; Sengupta, P.; McDonald, R. D.; Lancaster, T.; Blundell, S. J.; Pratt, F. L.; Cox, S.; Harrison, N.; Manson, J. L.; Southerland, H. I.; Schlueter, J. A. *New J. Phys.* **2008**, *10*, 083025.
- (12) e.g. Manson, J. L.; Schlueter, J. A.; Funk, K. A.; Southerland, H. I.; Twamley, B.; Lancaster, T.; Blundell, S. J.; Baker, P. J.; Pratt, F. L.; Singleton, J.; McDonald, R. D.; Goddard, P. A.; Sengupta, P.; Batista, C. D.; Ding, L.; Lee, C.; Whangbo, M.-H.; Franke, I.; Cox, S.; Baines, C.; Trial, D. *J. Am. Chem. Soc.* **2009**, *131*, 6733.
- (13) Blundell, S. J. *Contemp. Phys.* **1999**, *40*, 175.
- (14) Steele, A. J.; Lancaster, T.; Blundell, S. J.; Baker, P. J.; Pratt, F. L.; Baines, C.; Conner, M. M.; Southerland, H. I.; Manson, J. L.; Schlueter, J. A. *Phys. Rev. B* **2011**, *84*, 064412.
- (15) Yasuda, C.; Todo, S.; Hukushima, K.; Alet, F.; Keller, M.; Troyer, M.; Takayama, H. *Phys. Rev. Lett.* **2005**, *94*, 217201.